Dichromatometry

(pages 104-105)

1. Calculate the amount (m/g) of K$_2$Cr$_2$O$_7$ that we need to prepare 250ml of titrant solution with concentration 0,01 mol.dm$^{-3}$.

2. Calculate concentration of titrant solution that we prepared from …X..g of K$_2$Cr$_2$O$_7$ weighed on analytical balance and dissolved in 250 ml of water in volumetric flask.

3. We weighed …Y g... of the sample on analytical balance, dissolved it, transferred it quantitatively to 100 ml volumetric flask and added sufficient amount of acid. We pipetted 20 ml of that solution and titrated it with K$_2$Cr$_2$O$_7$ titrant solution (from task 2) using diphenylamine indicator. Average consumption from 3 parallel titrations was ..Z ml…of K$_2$Cr$_2$O$_7$ titrant solution. Calculate Fe$^{2+}$ content % in the sample.